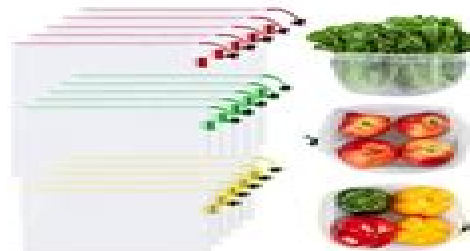


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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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 $8.3 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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 $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles } \text{O}_2$

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 $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 b determine the theoretical yield of Al_2O_3
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 150.1 \text{ g Al}_2\text{O}_3$
 $302.3 \text{ g Al}_2\text{O}_3 \times \frac{1 \text{ mol Al}_2\text{O}_3}{269.82 \text{ g/mol}} \times \frac{4 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} = 1.81 \text{ mol Al}$
 $1.81 \text{ mol Al} \times 27 \text{ g/mol} = 48.9 \text{ g Al}$
 c determine the percent yield
 $\frac{48.9 \text{ g Al}}{100.260 \text{ g}} \times 100 = 48.8\%$
 theoretical 302.3 g

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