Chemistry - Topic 15 - Transition Metals

TRANSITION METAL & elements that form one or more stable ions with incompletely filled O-craited MGANO & a species with one or more ione paid of electrons that surrounds a transition metal ion using coordinate bonding to form a complex ion

THE 3d SERIES OF THE PERIODIC TABLE

c) you need to unow now to work out the electronic configuration for any atom or ion.

SCANDIUM and ZING only form one ion such and their ions have as incomplete Dark cals.

GENERAL IDEAS ABOUT TRANSITION METALS

- * They are hard solids with high meeting and balling temperatures
- * Can have variable acidation numbers so they make good catalysts
- A They form coloured ions when in solution as they absorb some lights.

REMEMBER: Copper and Chromium discovey the author principle

ritis simply more atable por the electrons to pill like this .

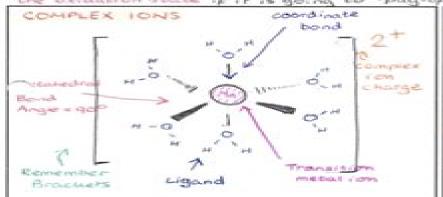
REMEMBER: Electrons from the 4s orbital are lost before those in 3d !

OTHER 10N3 V3" = 45° 3d° Mn2" = 45° 3d° Fe 3" = 45° 3d° Fe 3" = 45° 3d°

N:2+ = 450 3d2

WHY DO TRANSITION METALS HAVE VARIABLE OXIDATION STATES?

The compound pormed is the one where the most energy is released as the more energy that is released, the more scance the compound becomes. You need lots of energy to concer the metal but then more energy is released in hydration or lattice permation. As most of the electrons being removed in bransition metals, are in the Oranbibal, it is not too costly to increase the availation state if it is going to "pay-off" during lattice permation.



COMMON	MORE INFORMATION
Water	There are how lone pour of
HzO	Water often forms occuhedrons
Ammonia, NH ₃	One long pour of executions on the nutring on a both Ammonia is small so forms occahedral structures.
Hydrorde Ion OH	Forms hydroxo complexes, charge on the form was reduce the charge of the complex form.
Chloride rom	Circuit one large so opten cannot form actahulral completes so form hetrohedral completes

HORE ON VARIABLE OXIDATION STATES

Transition metals can donate and recieve electrons and are able to act as acidising agents and as brank reducing agents. The sub-shells of the D-orbital can easily gain or lose electrons meaning the energy dispersences between the acidation states is small.

- +2 state stability with respect to the 3+ state increases across a period.
- * Compounds with high acidation states hand to be acidising agencs eig. MnOst (Mn is 7")
 - compounds with low oxidation states hand to be reducing agents e.g. V2+, Fe2+ (2+and 2)

A2 Chemistry

Huangqi Zhang

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