



# Latest Work From Home Setup

**Edited by Rommel Sergio**



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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 mL of base is added section 17.3.2.1

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web solution a titrant volume 0 mL the solution ph is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then  $ph = -\log 0.100 = 1.000$   $ph = -\log 0.100 = 1.000$  b titrant volume 12.50 mL

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web expert answer 100 2 ratings the example is based upon henderson hasselbalch equation the equation gives relation between ph and pha of weak acid the equation is as  $ph = pka + \log \frac{[salt]}{[acid]}$  at equivalence point all a view the full answer transcribed image text

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web titration curves source openstax chemistry atoms first 2e strong acid strong base titrations when titrating a strong acid such as hydrochloric acid with sodium hydroxide you are reacting the hcl directly with naoh  $hcl(aq) + naoh(aq) \rightarrow h_2o(l) + nacl(aq)$

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web ab acid base titration curves purpose the purpose of this lab is to determine the shape of acid base titration curves theoretically and experimentally and to use an acid base titration to measure the effectiveness of antacid tablets summary a titration curve is the plot of the ph value of a solution vs the amount of titrant added

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4 6 4 6 5 9 5 9

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web strong acids and strong bases completely ionize in solution resulting in water and a salt for a weak acid there is only partial ionization the titration curve of a strong base strong acid showed a stretched out curve as it started with a slow gradual change in pH as it reached the equivalence point