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### SOME BASIC CON-CEPTS OF CHEMISTRY

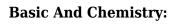
- Matter is substance which has mass and occupies space, e.g., water, air, box, table, etc.
- State of Matter. Solid, liquid and gaseous state represent three different states of mater. Gaseous mater at very high temperatures contains gaseous ions and free electron and is referred to as the plasma state, i.e., fourth state of matter.
- Homogenous. These have uniform composition throughout, e.g., elements and compounds. Mixture can also be homogeneous.
- Heterogeneous. These do not have uniform composition, e.g., mixture of iron fillings and sulphur powder.
- Element. It is a substance which cannot be further divided by chemical methods. It is a pure substance and consists of only one kind of atoms. They are metals, non-metals and metalloids (resemble both metal and non-metal), e.g., F. Al, Ni, Au, Ag are metals. C. S. N<sub>2</sub>, I<sub>2</sub> are non-metals. B. Si, As, Sb are metalloids. They are pure substances.
- Compound. It is a pure substance made up of two or more elements combined together in fixed ratio, e.g., NaCl. FeS, CuSO<sub>4</sub>, H<sub>2</sub>O, etc. Its components cannot be separated by simple physical methods.
  - (a) Organic compounds. They are compounds containing carbon and few other elements like H, N, S, halogens, phosphorus, oxygen etc. e.g., sugar, urea, glucose, etc.
  - (b) Inorganic compounds. They are made up of two or more elements and obtained form

minerals and ores, e.g., NaCl, KCl, CaCO<sub>4</sub>, etc.

- Mixture. It consists of two or more elements or compounds in any ratio. Its components can be separated by simple physical methods. It can be homogeneous (e.g., solution of salt in water, sugar in water). It can be heterogeneous, e.g., smoke dust particles in air, air, etc. Homogeneous mixture is also called solution.
- Separation of Mixtures. Mixtures are separated into its components depending upon the property of components in which they differ.
  - (a) Filtration. It is used to separate suspended (insoluble) component from a soluble component which passes into the filtrate and the insoluble component is left behind on the filter paper. Soluble component can be obtained by evaporation.
  - (b) Simple Distillation. It is used for separation of miscible liquids which differ in their boiling points appreciably, e.g., benzene and toluene can be separated by simple distillation.
  - (c) Fractional Distillation. It is used for separation of miscible fiquids which differ in their boiling points appreciably, e.g., benzene and toluene can be separated by simple distillation.
  - Molecules. They are identifiable units of matter consisting of two or more atoms of the same element or of different elements combined in a definite ratio.
  - Atomicity. It is defined as the number of atoms present in a molecule.
  - Monoatomic. They consist of only one

## **Basic And Chemistry**

Jessica J Manson



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