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web google classroom for the following redox equation $\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$ 2aq $\text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O}$ 4aq what will be the stoichiometric coefficient of H_2SO_4 if the above equation is balanced in an acidic medium stuck use a hint report a problem

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stages first bromine reacts with OH^- ions to form bromide ions and BrO^- ions second the BrO^- ions react to form bromide ions and bromate VO_3^- ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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